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2 Marks Questions1.The boiling point of hydrocarbons decreases with increase in branching. Give reason. Ans. Branching result into a more compact (nearly spherical) structure. This reduces the effective surface area and hence the strength of the Vander wall's forces, thereby leading to a decrease in the boiling point.2.Unsaturated compounds undergo addition reactions. Why?Ans. Unsaturated hydrocarbon compounds contain carbon - carbon double or triple bonds. The π -bond is multiple bond is unstable and therefore addition takes place across the multiple bonds.3.To which category of compounds does cyclohexane belong?Ans. Saturated alicyclic hydrocarbons.4.Draw the structure of the following compounds all showing C and H atoms. (a) 2-methyl -3-iso propyl heptanes (b) Dicyclopropyl methane.Ans.(a)(b)(dicyclopropyle methane)5.Draw all the possible structural isomers with the molecular formula C_6H_{14} . Name them.[2.5]Ans. (i) $CH_3 - CH_2 - CH_2 - CH_2 - CH_2 - CH_3$ (n- hexane)6.Sodium salt of which acid will be needed for the preparation of propane? Write chemical equation for the reaction.Ans. Butanoic acid, $CH_3CH_2CH_2COO-Na+NaOH \rightarrow CH_3CH_2CH_3+Na_2CO_3$.7.Cyclobutane is less reactive than cyclopropane. Justify. Ans . In cyclobutane molecule, the C-C-C bond angle is 90 while it is 600 in cyclopropane. This shows that the deviation from the tetrahedral bond angle (1090) in cyclobutane is less than in cyclopropane. In other words, cyclopropane is under great strain compared with cyclobutane and is therefore more reactive.8.How will you prepare isobutane?Ans . Isobutane is obtained by decarboxylation of 3-methyl butanoic acid with soda lime at 630K.9.The boiling point of alkanes shows a steady increase with increase in molecular mass. Why? Ans. This is due to the fact that the intermolecular van der Waals forces increase with increase of the molecular size or the surface area of the molecule.10.Pentane has three isomers i.e; pentane, 2-methyl butane and 2,2-dimethyl propane . The b.p of pentane is 309.1K whereas 2,2-dimethyl propane shows a b.p of 282.5k. Why? Ans .With the increase in number of branched chains, the molecule attains the shape of a sphere. This results in smaller area of contact and therefore weak inter molecular forest between spherical molecules, which are overcome a relatively lower temperatures.11.Draw the New man's projection formula of the staggered form of 1,2-dichloro ethane.Ans. 12.All the four C-H bonds in methane are identical. Give reasons. Ans .The four C-H bonds of methane are identical because all of these are formed by the overlapping of the same type of orbital's i.e; hybrid orbital's of carbon and s-orbital of hydrogen.13.When alkanes are heated, the C-C bonds rather than the C-H bonds break. Give reason. Ans. When alkanes are heated, the C-C bonds rather than the C-H bonds breaks because the C-C bond has a lower bond energy ($\Delta H=83K \text{ Cal/mole}$) than the C-H bond ($\Delta H=99 \text{ K Cal / mole}$).14.How would you convert cyclohexane to benzene?Ans . Cyclohexane when treated with iron or quartz in a red hot tube undergoes oxidation to form benzene.15.OEHow is iso-butane prepared? Ans .By decarboxylation of 3 - methyl butanoic acid with soda lime at 630 K.16.Why the addition of inert gas does not change the equilibrium?Ans. It is because the addition of an inert gas at constant volume does not change the partial pressures or the molar concentrations of the substance involved in the reaction.17.The equilibrium constant of a reaction increases with rise in temperature. Is the reaction exo - or endothermic?Ans. The equilibrium constant increases with a rise in temperature. Therefore, the reaction is endothermic.18.Using Le - chatelier principle, predict the effect of (a) decreasing the temperature (b) increasing the temperature in each of the following equilibrium systems: (i) $N_2 (g) + 3H_2(g) \rightleftharpoons 2NH_3(g) + (ii) N_2(g) + O_2(g) \rightleftharpoons 2NO(g)$ Ans.(i) For an exothermic reaction increase in temperature shifts the equilibrium to the left and decrease in temperature shifts it to the left.(ii) For an endothermic reaction increase in temperature shifts the equilibrium to the right and decrease in temperature shifts it to the right.19.(i) In the reaction equilibrium $A + B \rightleftharpoons C + D$. What will happen to the concentrations of A, B and D if concentration of C is increased.(ii) what will happen if concentration of A is increased?Ans. (i) For an equilibrium reaction $A + B \rightleftharpoons C + D$ if the concentration of a product is increased, the concentration of other components changes in such a way that the conc of C decreases and vice - versa.If the conc of C is increased the conc of D will decrease and those of A and B will increase simultaneously so that the numerical value of Kc is the same and vice - versa. The equilibrium shifts to the left.(ii) If the conc of A is increase, conc of B will decrease and those of C and D will increase simultaneously so that the numerical value of Kc is the same and vice - versa. The equilibrium shifts to the right.20.How is alkene produced by Kolbe's electrolytic method?Ans.21.How is alkene prepared from alcohol by acidic dehydration?Ans .Alcohols on heating with concentrated sulphuric acid form alkenes with the elimination of one water molecule.22.How are trans alkenes formed by alkynes? Ans. Alkynes on reduction with sodium in liquid ammonia form trans alkenes.23.How are cis - alkenes formed by alkynes?Ans. Alknes on partial reduction with calculated amount of dihydrogen in the presence of palladised charcoal partially deactivated with poisons like sulphur compounds or quinoline give cis-alkene.24.Stale Markownikov's Rule.Ans. It states that when a polar compound is added to an unsymmetrical alkenes, or alkynes positive part goes to the most substituted carbon atom and negative part goes to the least substituted carbon atom.25.Write the chemical equations of reactions involved in ozonolysis of alkenes. Ans. It is a process in which alkenes react with ozone to form ozonide which on reduction in presence of Zn give aldehyde and ketones. E.g;26.How will you distinguish between butene - 1 and butene - 2?Ans. Butene - 1 and butene - 2 can be distinguished either by ozonolysis or by oxidation with acidic $KMnO_4$ solution which they give different carbonyl compounds.27.State kharasch effect.Ans.It states that in presence of peroxides such as benzoyl peroxide, addition of HBr (but not of HCl or HI) to unsymmetrical alkenes occurs contrary to Markontkov's rule. $CH_3CH = CH_2 + HBr \rightarrow CH_3 - CH_2 - CH_2 - Br$ (1 - bromopropane)28.How is alkyne prepared from calcium carbide? Ans. Calcium carbide is treated with water to get ethyne. $CaC_2 + 2H_2O \rightarrow Ca(OH)_2 + C_2H_2$ 29.How is alkyne prepared by Kolbe's method?Ans.30.How is alkyne prepared from vicinal dihalides? Ans. Vicinal dihalides on treatment with alcoholic potassium hydroxide undergo dehydrohalogenation. One molecule of hydrogen halide is eliminated to form alkenyl halide which on treatment with sodiumamide gives alkynes.31.How will you distinguish between ethylene and methane? Ans. Ethylene discharges bromine water colour and Baeyer's reagent colour while methane does not.32.Although acetylene is acidic in nature, it does not react with NaOH or KOH. Give reason? Ans. Acetylene is a very weak acid ($pK_a=25$) and hence only an extremely strong base like amide ion (NH_2^-) can successfully remove a proton.33.Write the conversion of ethene to ethyne. Ans.34.How would you distinguish between butyne - 1 and butyne - 2?Ans. Butyne - 1 ($CH_3CH_2C \equiv CH$), having an acetylene hydrogen atom will give white precipitate with ammonical silver nitrate and red precipitate with ammonical cuprous chloride. On the other hand, butyne - 2 ($CH_3C \equiv CCH_3$) having no acetylene hydrogen atom does not respond to either of the two reagent.35.How would you carry out the following conversion propene to ethyne. Ans. 36.How will you convert propyne to propanone? Ans.37.How will you convert ethyne to ethane? Ans.38.Convert 2- butyne to trans - 2- butane. Ans.39.How will you prepare 3-methyl but -1 - yne by starting with ethyne?Ans.40.Write the IUPAC name of the following compound- Ans.(i) 4 - phenyl - but - 1 - ene.(ii) 2 - Methyl phenol.41.What do you mean by delocalization? Ans.Delocalisation - Delocalisation implies that pairs of bonding electrons extend over three or more atoms and belong to the whole molecule. Delocalized π -orbitals are much larger than the localized π -orbitals and are therefore more stable.42.What do you understand by Resonance energy?Ans.The difference between the energy of the most stable contributing structure and the energy of the resonance hybrid is known as resonance energy. In case of benzene, the resonance hybrid has (147KJ/mol-1) less energy than either A to B. Thus resonance energy of benzene is 147KJ/mole.43.How is phenol reduced to benzene?Ans.44.How is aromaticity of a compound judged? Ans.The following characteristics decides aromaticity of a compound:-(i) Planarity(ii) Complete delocalization of the π -electrons in the ring.(iii) Presence of $(4n+2)$ π electrons in the ring where n is an integer ($n=0, 1, 2, \dots$)This is often referred to as Huckel Rule.45.Give some examples of aromatic compounds.Ans.46.How will you account for the structure of benzene?Ans. All the six carbon atoms in benzene are sp^2 hybridised. Two sp^2 hybrid orbitals of each carbon atom overlap with sp^2 hybrid orbitals of adjacent carbon atoms to form six C-C sigma bonds with are in the hexagonal plane. The remaining sp^2 hybrid orbital of each carbon atom overlaps with s-orbital of a hydrogen atom to form six C-H sigma bonds. Each carbon atom is now left with one hybridized p-orbital perpendicular to the plane of the ring.The unhybridized p-orbital of C-atoms are close enough to form a bond by lateral overlap.47.How is benzene prepared from aromatic acids? Ans. Sodium salt of benzoic acid on heating with soda lime gives benzene.48.How is phenol reduced to benzene?Ans. Phenol is reduced to benzene by passing its vapours over heated zinc dust.49.Why is benzene extra ordinarily stable though it contains three double bounds? Ans. Due to resonance.50.What is friedel craft's reaction? Give an example.Ans. When benzene or its derivative reacts with alkyl halide in presence of $AlCl_3$, we get alkyl benzene.51.What happens when benzene is oxidized at 770K in presence of V_2O_5 ? Give chemical equation. Ans.52.How will you convert benzene to iodobenzene? Give chemical equation. Ans.53.What are electrophilic substitution reactions? Ans. Those reactions in which weaker electrophile are replaced by a stronger electrophile are called electrophilic substitution reactions.54.How will you distinguish between Ethene and benzene.Ans. Ethene discharges bromine water colour and Baeyer's reagent colour while benzene does not.55.How is benzene converted to benzene hexachloride? Ans. Under ultra-violet light, three chlorine molecules add to benzene to produce benzene hexachloride, $C_6H_6Cl_6$ which is also called gammaxane.56.How will you convert benzene to hexachlorobenzene? Ans. Benzene on treatment with of chlorine in the presence of anhydrous $AlCl_3$ in dark yields hexachloroben - zene (C_6Cl_6)

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